

(19)



Europäisches Patentamt
European Patent Office
Office européen des brevets



(11) Publication number:

0 601 594 A1

(12)

EUROPEAN PATENT APPLICATION

(21) Application number: 93119965.7

(61) Int. Cl. 5: **B01J 13/04**

(22) Date of filing: 10.12.93

(30) Priority: 11.12.92 JP 353353/92
16.03.93 JP 81544/93

(43) Date of publication of application:
15.06.94 Bulletin 94/24

(84) Designated Contracting States:
DE FR GB IT

(71) Applicant: **ASAHI GLASS COMPANY LTD.**
1-2, Marunouchi 2-chome
Chiyoda-ku Tokyo 100(JP)

(72) Inventor: **Sunahara, Kazuo, Asahi Glass
Company Ltd.**
Chuo Kenkyusho,
1150, Hazawa-cho,
Kanagawa-ku
Yokohama-shi, Kanagawa-ken(JP)

(74) Representative: **Müller-Boré & Partner**
Patentanwälte
Isartorplatz 6
D-80331 München (DE)

(54) **Process for producing crystalline microballoons.**

(57) A process for producing crystalline microballoons, which comprises atomizing a solution or dispersion having an inorganic material or its precursor dissolved or dispersed in a liquid medium, to form droplets of the solution or dispersion, supplying the droplets into a high temperature atmosphere adjusted so that the liquid medium would rapidly evaporate and the inorganic material or its precursor would sinter or fuse to form crystalline microballoons of the inorganic material, and recovering the microballoons thereby formed.

EP 0 601 594 A1

The present invention relates to a process for producing crystalline microballoons and new crystalline microballoons produced thereby.

Heretofore, as microballoons (microbubbles or hollow microspheres) made of an inorganic material, vitreous fine hollow bodies so-called glass microballoons or microballoons made of, e.g., alumina or zirconia, have been known. Further, resinous fine hollow bodies so-called plastic microballoons have been known. These microballoons are widely used as fillers in solid materials such as resins for weight reduction or heat insulation for transporting equipments, building materials, paints or refractories to meet various demands of the age.

For production of glass microballoons, a method is known which comprises heat-melting glass at a high temperature and blowing it off in the form of particles while foaming it with a blowing agent, to form microballoons (U.S. Pat. No. 3,365,315, No. 4,391,646, No. 4,767,726 and No. 5,064,784). Likewise, plastic balloons are produced by a method of heating, melting and foaming.

In these methods, it is necessary to heat-melt the material for microballoons itself. Accordingly, the material for microballoons has been limited to the glass which can be easily melted at a commercially practical temperature, for example, at a temperature of not higher than 1500 °C. Thus, the material for glass microballoons has been limited to the one having a melting point which is not so high, such as soda-lime glass or soda borosilicate glass.

Further, in such a conventional method for glass microballoons, microballoons are produced by such an operation that molten glass is blown off in the form of particles while being foamed, as mentioned above, whereby the average particle size of the resulting microballoons is usually as large as from 50 to 5000 µm, and it has been difficult to obtain balloons having a very small average particle size.

For production of ceramic microballoons, a method similar to that of producing glass microballoons is known. In these methods, it is necessary to heat-melt the material for microballoons. Thus, an alkali component is added to the materials in order to decrease the melting temperature.

Furthermore, in the conventional method involving heating and melting of the material, the apparatus is necessarily large-sized and expensive, and such a method is not suitable for producing a small quantity of various types of microballoons.

It is an object of the present invention to provide a novel process for producing crystalline microballoons, whereby a room for selection of the material for crystalline microballoons is wide, it is possible to readily obtain microballoons made of a material which has never been used before in a conventional material with an extremely small average particle size, and the apparatus may be small-sized and inexpensive.

Thus, the present invention provides a process for producing crystalline microballoons, which comprises atomizing a solution or dispersion having an inorganic material or its precursor dissolved or dispersed in a liquid medium, to form droplets of the solution or dispersion, supplying the droplets into a high temperature atmosphere adjusted so that the liquid medium would rapidly evaporate and the inorganic material or its precursor would fuse or preferably sinter to form microballoons of the inorganic material, and recovering the microballoons thereby formed.

In the accompanying drawings;

Figure 1 is a schematic view illustrating an embodiment of the apparatus useful for carrying out the process of the present invention.

Figure 2 is an electron microscopic photograph showing the structures of crystalline microballoons obtained in Example 3.1.

Now, the present invention will be described in detail with reference to the preferred embodiments.

The inorganic material to be dissolved or dispersed in a liquid medium in the process of the present invention, may be selected from an extremely wide range of inorganic materials. When an inorganic material desired for crystalline microballoons is used by itself as dissolved or dispersed in a liquid medium, it may, for example, be an oxide such as silica, alumina, spinel, mullite, cordierite, titanium oxide or tin oxide, a nitride such as silicon nitride, aluminum nitride, titanium nitride, zirconium nitride or boron nitride, a carbide such as silicon carbide, titanium carbide or boron carbide, a metal such as copper, aluminum or lead, or carbon.

When a precursor which forms an inorganic material for crystalline microballoons by a reaction in the condition of preparation of the microballoons of the present invention, is used as dissolved or dispersed in a liquid medium, one precursor may be used, or a plurality of precursors may be used as a mixture. Specifically, the precursor may be various elemental single substances, various acid salts such as sulfates, nitrates, phosphates, acetates and oxalates, hydroxides, chlorides, sulfides, oxides, nitrides, carbides, cyanides, and chelate compounds. As a typical example of the liquid medium, water is preferably used. Otherwise, an organic medium such as a halogenated hydrocarbon, an ether, an alcohol, a ketone, a

hydrocarbon or an organic acid, may be used as the liquid medium. From the viewpoint of handling efficiency, the one having a boiling point of from 50 to 200 °C, particularly from 80 to 120 °C, is preferred.

The concentration in the solution or the dispersed concentration and particle size in the dispersion of the inorganic material or its precursor affect the particle size, specific gravity, strength, etc. of microballoons thereby produced. The concentration of the solution or dispersion is preferably from 0.1 to 80 wt%, more preferably from 1 to 10 wt%.

On the other hand, the particle size of the inorganic material or its precursor in the dispersion is preferably from 1 to 1000 nm, more preferably from 1 to 100 nm. The concentration is preferably from 0.1 to 50 wt%, more preferably from 1 to 50 wt%. The dispersion is preferably a uniform dispersion. Otherwise, it may be formed into a uniform colloidal solution by emulsifying it using a suitable emulsifier, as the case requires.

The particle size, specific gravity, strength, etc. of the resulting crystalline microballoons can be controlled by incorporating a suitable additive to the above solution or dispersion, as the case requires. Such an additive is preferably used in amount of from 0.1 to 10 wt% based on the material for microballoons. As such an additive, a material may be used which has a melting point lower than the melting point of the material for microballoons and which is capable of controlling the growing rate of the crystals for forming microballoons. For example, when alumina microballoons are to be prepared, silica, magnesia, forsterite or various types of clay having a melting point lower than alumina and being capable of controlling the growth of alumina crystals, may, for example, be used.

In the process of the present invention, the above solution or dispersion is firstly atomized into the droplets. The atomizing method is not particularly limited. However, a conventional method such as a supersonic method, a spray method or a rotary method may preferably be employed. The particle size of the droplets affects the particle size of the microballoons to be produced. The average particle size is preferably from 0.1 to 1000 µm, more preferably from 10 to 100 µm, although it depends also on the types of the inorganic material and the liquid medium.

The droplets are then supplied into a high temperature atmosphere. Here, the temperature and the atmosphere are controlled. The temperature is required to be within a temperature range wherein the liquid medium used will rapidly evaporate, and the inorganic material or its precursor will sinter or fuse to form crystalline microballoons of the inorganic material. The temperature for rapid evaporation is preferably at least three times, more preferably from 5 to 20 times, higher than the boiling point of the liquid medium, as represented by centigrade (°C). When water is used as the liquid medium, the temperature for rapid evaporation is preferably from 300 to 2200 °C.

In the process of the present invention, if the material is heated beyond its melting point, adjacent crystal grains will fuse to one another and abnormally grow, whereby the strength of the resulting microballoons will be low. Accordingly, the high temperature atmosphere is preferably at a temperature not higher than the melting point of the material for microballoons, more preferably lower by about 100 °C, particularly by about 200 °C, than the melting point of the inorganic material.

The high temperature atmosphere may be set up in a single step system wherein the temperature is maintained uniformly throughout the operation, or in a two step system wherein the temperature is changed from a temperature range of a first step wherein the liquid medium will rapidly evaporate to a temperature range of a second step wherein the inorganic material or its precursor will fuse or melt to form microballoons, or may be set up in a multistep system of more than two steps. For example, in a case where the liquid medium is water and the inorganic material is alumina, the high temperature atmosphere may be adjusted so that the temperature in the vicinity of the inlet is preferably from 500 to 1000 °C and the temperature in the vicinity of the outlet is preferably from 1000 to 1800 °C.

It is a characteristic of the present invention that the inorganic material for crystalline microballoons is not necessarily heated to its melting point, so long as it is heated to a sintering temperature to obtain microballoons. By virtue of this characteristic, it is now possible to produce crystalline microballoons easily even from a material which used to be difficult to heat-melt. Further, it is possible to produce crystalline microballoons easily also from a material which contains no alkali component or a very small amount of an alkali component, which used to be difficult to melt.

The atmosphere for constituting the high temperature atmosphere is preferably an atmosphere suitable for the inorganic material and the liquid medium to be used. For example, in a case where crystalline microballoons made of a metal oxide are to be produced, an oxygen-containing atmosphere such as atmospheric air is selected for use. Likewise, in the case of a metal nitride, a nitrogen gas atmosphere is selected, and in the case of a metal, a reducing atmosphere such as hydrogen gas is selected. In the case where microballoons are produced from a precursor of an inorganic material by means of a reaction in the high temperature atmosphere producing microballoons of the present invention, an atmosphere suitable for

such a reaction is employed.

Spraying of the droplets into the high temperature atmosphere may be carried out by various means. The high temperature atmosphere may be formed in e.g. a tubular furnace or a fluidized bed furnace. Preferred specific means for spraying droplets include, for example, a supersonic atomizer, a spray atomizer and a rotary disc atomizer, and the droplets are sprayed by such an atomizer into the above furnace at a linear speed of the droplets being preferably at least 0.01 m/sec, more preferably from 0.1 to 10 m/sec.

The droplets are maintained in the above high temperature atmosphere usually from 10 seconds to 30 minutes, although the retention varies depending upon the type of the inorganic material, whereby microballoons will be formed, in some cases, accompanied by the above mentioned reaction. The formed crystalline microballoons will be collected in such a manner that, for example, in the case of a tubular furnace, microballoons discharged from the tubular furnace are collected by means of a bag filter or a liquid medium such as water. The collected crystalline microballoons may have various average particle sizes depending upon the sizes of the droplets.

As an apparatus to be used for carrying out the process of the present invention, an apparatus having the construction as shown in Figure 1, may, for example, be employed. In Figure 1, the solution or dispersion 1 is atomized by an atomizer 2 to form fine droplets, which are then supplied to a tubular furnace 3. The tubular furnace comprises a reaction tube 4 and a heater 5. The atmosphere is heated to a predetermined temperature by the heater 5. The droplets will be transported in the reaction tube 4 towards the right hand side in Figure 1 by a gas stream created by a volume expansion accompanied by evaporation of the liquid medium of the droplets. Here, a transporting gas may separately be introduced. Microballoons formed in the reaction tube may be recovered by various methods. In Figure 1, microballoons are introduced through a pipe 6 equipped with a heater to prevent condensation into a recovery filter 7 and recovered therein. For the recovery filter, a suction apparatus 8 may be employed to increase the recovery efficiency.

According to the process of the present invention, it is possible to obtain spherical crystalline microballoons having an average particle size of from 0.1 to 300 μm , if desired, substantially perfectly spherical microballoons having an average particle size of from 0.1 to 100 μm . Especially, it is possible to obtain fine microballoons with an average particle size of not more than 50 μm , which have been difficult to obtain heretofore. For example, it is possible to obtain microballoons with an average particle size of from 1 to 20 μm , particularly from 1 to 10 μm . The wall thickness of the microballoons can be controlled by adjusting e.g. the concentration or particle size of atomized droplets or the heating conditions. It is possible to obtain microballoons having a bulk density of from 0.01 to 2.0 g/cm^3 , although the bulk density may vary depending upon the material. Further, since the microballoons produced by the present invention have a crystalline structure, preferably a polycrystalline network structure and have substantially perfectly spherical shapes, the pressure collapse strength is as high as at least 1000 kg/cm^2 , especially at least 1500 kg/cm^2 . Further, the flowability is excellent, and when the microballoons are mixed with a resin or the like, they are free from breakage during the mixing, and the surface smoothness of the resulting molded product of the resin will be excellent.

The alkali content of the microballoons obtainable by the present invention can be controlled within a wide range. If necessary, it is possible to obtain crystalline microballoons having a very low alkali content as never has been possible heretofore. It is possible to produce microballoons having excellent water resistance, preferably with an alkali elution degree of not higher than 0.01 meq/g, more preferably not higher than 0.001 meq/g. Here, the alkali elution degree is determined in such a manner that a sample is put in pure water so that it constitutes 10 wt% and left to stand at 60°C for 24 hours, and then the alkali elution degree is measured. In the process of the present invention, the presence of an alkali component is not required for formation of microballoons by sintering, as the melting step is not required. The alkali content in crystalline microballoons is preferably not higher than 0.1 wt%, more preferably not higher than 0.01 wt%.

According to the process of the present invention, it is possible to produce crystalline microballoons of various metal oxides. For example, it is possible to produce crystalline microballoons made of an oxide of a transition metal such as scandium, titanium, vanadium, chromium, manganese, iron, cobalt, nickel, copper, zinc, yttrium, zirconium, niobium, molybdenum, ruthenium, rhodium, palladium, tantalum, tungsten, lanthanide or actinide, or an oxide of a semi metal element such as boron, aluminum, silicon, gallium, germanium, indium, tin, antimony, thallium, lead or bismuth. It is also possible to produce crystalline microballoons made of an oxide of a Group 2A element (such as beryllium, magnesium, calcium, strontium or barium) or a Group 1A element (lithium, sodium or potassium). Among them, those which are instable by themselves, such as oxides of Group 1A elements, may preferably be used as chemically reacted with

suitable other components.

Crystalline microballoons made of a double oxide may likewise be produced. Specific examples of such a double oxide include zinc ferrite ($\text{ZnO} \cdot \text{Fe}_2\text{O}_3$), barium ferrite ($\text{BaO} \cdot 6\text{Fe}_2\text{O}_3$), aluminum titanate ($\text{Al}_2\text{O}_3 \cdot \text{TiO}_2$), barium titanate ($\text{BaO} \cdot \text{TiO}_2$), mullite ($3\text{Al}_2\text{O}_3 \cdot 2\text{SiO}_2$), cordierite ($2\text{MgO} \cdot 2\text{Al}_2\text{O}_3 \cdot 5\text{SiO}_2$), spinel ($\text{MgO} \cdot \text{Al}_2\text{O}_3$), steatite ($\text{MgO} \cdot \text{SiO}_2$), zircon ($\text{ZrO}_2 \cdot \text{SiO}_2$), forsterite ($2\text{MgO} \cdot \text{SiO}_2$), eucryptite ($\text{LiO}_2 \cdot \text{Al}_2\text{O}_3 \cdot 2\text{SiO}_2$), β -spodumen ($\text{LiO}_2 \cdot \text{Al}_2\text{O}_3 \cdot 4\text{SiO}_2$) and aluminum borate ($9\text{Al}_2\text{O}_3 \cdot 2\text{B}_2\text{O}_3$).

According to the process of the present invention, the material for crystalline microballoons is not limited to an oxide, and it is possible to obtain crystalline microballoons made of a material such as a nitride or carbide, which have not been obtained heretofore. As the nitride, aluminum nitride (AlN; wurtzite structure or hexagonal system), silicon nitride (Si_3N_4 ; hexagonal system), tantalum nitride (TaN; hexagonal system) or boron nitride (BN; hexagonal system) may, for example, be preferred. As the carbide, silicon carbide (SiC; diamond structure), zirconium carbide (ZrC; isometric system) or tungsten carbide (WC; hexagonal system) may, for example, be preferred.

Likewise, it is possible to produce crystalline microballoons made of a metal. The metal is not particularly limited, and various metal elements may be used alone or in combination as an alloy or intermetallic compound, or as a mixture of two or more of them. The metal is preferably the one which is hardly reactive to form a compound such as an oxide. For example, a noble metal such as gold, silver, copper or platinum, or lead, is preferred. The metal may be dispersed as a simple substance in the droplets, or a metal may be formed by a reaction at a high temperature. It is also possible to produce crystalline microballoons of a simple substance such as graphite.

The mechanism for formation of crystalline microballoons in the present invention is not fully understood, but is considered to be as follows. When the solution or dispersion of a material is atomized and supplied in the form of droplets into the high temperature atmosphere, the liquid medium will rapidly evaporate at the surface of the droplets. As a result, in the case of the solution, the material dissolved in the medium will precipitate under super saturated condition along the interface of droplets in spherical forms corresponding to the shapes of the droplets. Likewise, in the case of the dispersion, the material dispersed in the medium will aggregate along the interface of the droplets in spherical forms corresponding to the shapes of the droplets.

The liquid medium remaining in the interior of droplets will evaporate and dissipate through spaces of precipitates precipitated in spherical forms into the atmosphere, and at the same time, the material dissolved or dispersed in the interior of the droplets will move in the centrifugal direction along with the evaporation of the liquid medium and will precipitate around the above mentioned spherical precipitates, so that the precipitates will grow and densify and the interior of the spherical forms will be hollow.

In the case where a precursor of an inorganic material is used, the desired inorganic material will be formed during the process of precipitation or aggregation, in some cases by a reaction with the atmosphere. For example, in the case of an oxygen-containing atmosphere, it may react with oxygen to form an oxide. Then, the precipitated material or the aggregated material is sintered or fused at a high temperature and further densified to finally form hollow microballoons having high strength.

Now, the present invention will be described in further detail with reference to Examples. However, it should be understood that the present invention is by no means restricted to such specific Examples.

Using various solutions or dispersions, microballoons were prepared by an apparatus as shown in Figure 1. In this apparatus, the solution or dispersion was atomized by a supersonic atomizer (frequency: 2MHz) into droplets, which were then introduced into a tubular furnace (length of the heating zone: 50 cm, diameter: 9 cm). Formed microballoons were collected by a bug filter (a glass cloth coated with a fluorine resin). In some Examples, a two step tubular furnace comprising two tubular furnaces of the same type disposed in series, was employed. The size of droplets and the temperature of the tubular furnace were suitably adjusted. The analysis of microballoons obtained in each Example was carried out by the following methods.

Shape: Microballoons were mixed with an epoxy resin and cured, the cured resin was cut, and the cut surface was polished to expose the cross sections of the microballoons. Then, gold was vapor-deposited thereon, and the shapes were inspected by a scanning electron microscope JSM-T300 model, manufacture by Nippon Denshi K.K.

Average particle size: Microballoons were fixed on an adhesive tape. Then, gold was vapor-deposited thereon, followed by observation by a scanning electron microscope JSM-T300 model manufactured by Nippon Denshi K.K., whereupon the average particle size (diameter) was calculated by the image analysis.

Crystal phase: Microballoons were pulverized for 30 minutes in an agate mortar, and the crystal phase was identified by an X-ray diffraction apparatus (Gaiga flex, trade name) manufactured by Kabushiki Kaisha Rigaku.

Bulk density: The bulk density was measured as a tap density of microballoons by a vibration type specific gravity measuring apparatus KRS-406 manufactured by Kuramochi Kagakukiki Seisakusho (measuring conditions: 1/3 Hz, up-down of 30 mm: 700 times).

True specific gravity: The true specific gravity was measured by a gas-substitution method using an argon gas by means of a microvolume picnometer manufactured by Shimadzu Corporation. Here, the true specific gravity is the one obtained by dividing the mass of microballoons by the volume including the space portions.

Pressure collapse strength: The pressure at the time when 10% of microballoons were broken under hydraulic pressure, was obtained by a hydraulic pressure resistance tester, manufactured by Nichion Irikakikai Seisakusho.

Alkali elution degree: A sample was put in pure water so that it constituted 10 wt% and left to stand at 60 °C for 24 hours, whereupon the amount of alkali metal elements eluted into pure water was measured by an elemental analysis by a plasma emission spectrochemical analyzer ICPS-1000 model manufactured by Shimadzu Corporation. The detection limit of alkali metals was 0.1 ppm.

EXAMPLE 1.1

A dispersion having 10 wt% of manganese oxide particles having a particle size of 2 nm dispersed in water, was prepared. This dispersion was atomized at a flow rate of 1 ml/min to form droplets having an average particle size of 30 μ m, which were introduced into a tubular furnace wherein atmospheric air was maintained at 1600 °C. The obtained microballoons were spherical polycrystals of manganese oxide (tetragonal system MnO_2) having an average particle size of 1.5 nm, a bulk density of 0.032 g/cm³, a true specific gravity of 0.50 g/cm³ and a pressure collapse strength of 1430 kg/cm². No alkali elution was observed.

EXAMPLES 1.2 TO 1.12

Using the dispersions as identified in Table 1 as starting materials, microballoons were prepared in the same manner as in Example 1.1 except that the temperature of the tubular furnace was as shown in Table 1. The evaluation results are shown in Table 2. In each case, the microballoons were spherical polycrystals, and no alkali elution was observed.

Table 1

Example No.	Dispersed particles	Average particle size (nm)	Concentration (wt%)	Temp. (°C)
1.2	Zinc oxide	3	10	1400
1.3	Iron oxide	6	10	1400
1.4	Zirconium oxide (containing 5 mol % of yttrium oxide)	3	10	1600
1.5		3	10	1300
1.6		4	5	1750
1.7		3	10	1400
1.8		3	10	1600
1.9		3	10	1800
1.10		3	10	1800
1.11		3	10	1800
1.12		3	10	750

Table 2

Example No.	Crystal phase	Average size of microballoons (μm)	Bulk density (g/cm^3)	True specific gravity (g/cm^3)	Pressure collapse strength (kg/cm^2)
1.2	Wurtzite structure ZnO	1.5	0.022	0.48	1310
1.3	$\gamma\text{-Fe}_2\text{O}_3$	1.5	0.023	0.44	1660
1.4	Monoclinic cubic sustem Mixture $\text{ZrO}_2\text{-Y}_2\text{O}_3$	1.3	0.025	0.49	2930
1.5	Tetragonal system WO_2	1.6	0.080	0.96	1830
1.6	Tetragonal system CrO_2	1.3	0.030	0.49	1460
1.7	Rutile type RuO_2	1.9	0.029	0.46	1230
1.8	Cubic system CoO	1.3	0.049	0.66	1690
1.9	Cubic system Y_2O_3	1.3	0.023	0.43	1530
1.10	Hexagonal system La_2O_3	1.5	0.025	0.46	1480
1.11	Cubic system CeO_2	1.8	0.039	0.43	1530
1.12	Monoclinic system MoO_3	1.3	0.029	0.46	1090

55 EXAMPLE 2.1

Titanium tetrachloride was diluted with pure water to obtain a solution containing 5 wt% of TiCl_4 . This solution was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of 30

μm , which were introduced into a tubular furnace wherein atmospheric air was maintained at 1500°C . The obtained microballoons were spherical polycrystals of titanium oxide (rutile type TiO_2) having an average particle size of $1.6 \mu\text{m}$, a bulk density of 0.019 g/cm^3 and a true specific gravity of 0.22 g/cm^3 . No alkali elution was observed.

EXAMPLES 2.2 TO 2.6

Using the solutions as identified in Table 3 as starting materials, microballoons were prepared in the same manner as in Example 2.1 except that the temperature of the tubular furnace was as shown in Table 3. The evaluation results are shown in Table 4. In each case, the microballoons were spherical polycrystals, and no alkali elution was observed.

Table 3

Example No.	Solute	Concentration (wt%)	Temp. ($^\circ\text{C}$)
2.2	MnCl_2	5	1600
2.3	NiCl_2	5	1650
2.4	NbCl_5	1	1500
2.5	TaCl_5	1	1700
2.6	CuSO_4	1	1200

Table 4

Example No.	Crystal phase	Average Size (μm)	Bulk density (g/cm^3)	True specific gravity (g/cm^3)
2.2	Tetragonal system MnO_2	2.7	0.023	0.39
2.3	Trigonal system NiO	2.9	0.028	0.39
2.4	Monoclinic cubic system Nb_2O_5	2.7	0.020	0.37
2.5	Tetragonal system Ta_2O_5	2.7	0.031	0.57
2.6	$\text{Cu}_2\text{O-CuO}$ mixture	3.6	0.031	0.42

EXAMPLE 2.7

An aqueous solution containing 1 wt% of rhodium chloride (RhCl_3) was prepared. This solution was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of $30 \mu\text{m}$, which were introduced into a two step tubular furnace comprising a tubular furnace wherein atmospheric air was maintained at 1800°C and a tubular furnace wherein atmospheric air was maintained at 900°C , connected in series. The obtained microballoons were spherical polycrystals of rhodium oxide (trigonal system Rh_2O_3) having an average particle size of $3.6 \mu\text{m}$, a bulk density of 0.038 g/cm^3 and a true specific gravity of 0.54 g/cm^3 . No alkali elution was observed.

EXAMPLE 3.1

A colloidal solution having 10 wt% of aluminum oxide particles having a particle size of 6 nm dispersed therein, was prepared. This colloidal solution was atomized at a flow rate of 1 ml/min to form droplets having an average particle size of $90 \mu\text{m}$, which were introduced into a tubular furnace wherein atmospheric air was maintained at 1500°C . The resulting microballoons were spherical polycrystals of corundum ($\alpha\text{-Al}_2\text{O}_3$) having structures as shown in the electro microscopic photograph of Figure 2. In the photograph, the total length of the horizontal line at the bottom corresponds to $5 \mu\text{m}$. This photograph shows a region wherein fractured microballoons exist. Further, the microballoons had an average particle size of $3.5 \mu\text{m}$ and a true specific gravity of 0.33 g/cm^3 , and no alkali elution was observed.

EXAMPLE 3.2

Metal aluminum having an average particle size of 10 μm was dissolved in 50 wt% nitric acid to obtain a 2 wt% $\text{Al}(\text{NO}_3)_3$ aqueous solution. This solution was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of 90 μm , which were introduced into a tubular furnace wherein atmospheric air was maintained at 1600 °C. The resulting microballoons were spherical polycrystals of corundum. No alkali elution was observed. The evaluation results are shown in Table 5.

EXAMPLE 3.3

A 5 wt% AlCl_3 aqueous solution was atomized at a flow rate of 1 ml/min to form droplets having an average particle size of 88 μm , which were introduced into a tubular furnace wherein atmospheric air was maintained at 1500 °C. The obtained microballoons were spherical polycrystals of corundum. No alkali elution was observed. The evaluation results are shown in Table 5.

EXAMPLE 3.4

A colloidal solution having 10 wt% of silicon dioxide particles having a particle size of 5 nm dispersed in water, was prepared. This colloidal solution was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of 110 μm , which were introduced into a tubular furnace wherein atmospheric air was maintained at 1000 °C. The obtained microballoons were spherical polycrystals of α -quartz. No alkali elution was observed. The evaluation results are shown in Table 5.

EXAMPLE 3.5

A colloidal solution having 10 wt% of silicon dioxide powder having a particle size of 6 nm dispersed in water, was prepared. This colloidal solution was atomized at a flow rate of 1 ml/min to form droplets having an average particle size of 90 μm , which were introduced into a tubular furnace wherein atmospheric air was maintained at 1500 °C. The obtained microballoons were spherical polycrystals of cristobalite. No alkali elution was observed. The evaluation results are shown in Table 5.

Table 5

Example No.	Average size (μm)	Bulk density (g/cm^3)	True specific gravity (g/cm^3)	Pressure collapse strength (kg/cm^2)
3.2	6.3	0.034	0.34	2000
3.3	3.2	0.036	0.38	2100
3.4	6.3	0.030	0.41	1660
3.5	2.5	0.033	0.46	1560

EXAMPLE 4.1

A colloidal solution having 10 wt% of tin oxide powder having a particle size of 3 nm dispersed in water, was prepared. This colloidal solution was atomized at a flow rate of 1 ml/min to form droplets having an average particle size of 30 μm which were introduced into a tubular furnace wherein atmospheric air was maintained at 1200 °C. The obtained microballoons were spherical polycrystals of tin oxide (rutile type SnO_2) having an average particle size of 1.6 μm , a bulk density of 0.020 g/cm^3 , a true specific gravity of 0.46 g/cm^3 and a pressure collapse strength of 1330 kg/cm^2 . No alkali elution was observed.

EXAMPLE 4.2 to 4.7

Using the dispersions as identified in Table 6 as starting materials, microballoons were prepared in the same manner in Example 4.1 except that the temperature of the tubular furnace was as shown in Table 6. The evaluation results are shown in Table 7. In each case, the microballoons were spherical polycrystals, and no alkali elution was observed.

Table 6

Example No.	Dispersed particle	Particle size (nm)	Concentration (wt%)	Temp. (°C)
4.2	Germanium oxide	3	10	1200
4.3	Antimony oxide	3	10	1200
4.4	Bismuth oxide	3	10	750
4.5	Gallium oxide	3	10	1500
4.6	Beryllium oxide	4	10	1900
4.7	Magnesium oxide	4	10	1650

Table 7

Example No.	Crystal phase	Average size (μm)	Bulk density (g/cm ³)	True specific gravity (g/cm ³)	Pressure collapse strength (kg/cm ²)
4.2	Tetragonal GeO ₂	1.6	0.039	0.56	1590
4.3	Monoclinic system Sb ₂ O ₃	1.6	0.025	0.46	1230
4.4	Tetragonal Bi ₂ O ₃	1.6	0.40	0.86	1330
4.5	Monoclinic cubic system Ga ₂ O ₃	1.9	0.33	0.46	1330
4.6	Sodium-chloride type BeO	1.8	0.21	0.38	1760
4.7	Sodium-chloride type MgO	1.8	0.20	0.39	1360

EXAMPLE 5.1

Bismuth nitrate was diluted with pure water to obtain a solution containing 5 wt% of Bi(NO₃)₃. This solution was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of 30 μm, which were introduced into a tubular furnace wherein atmospheric air was maintained at 750 °C. Obtained microballoons were spherical polycrystals of bismuth oxide (tetragonal system Bi₂O₃) having an average particle size of 2.6 μm, a bulk density of 0.039 g/cm³ and a true specific gravity of 0.69 g/cm³. No alkali elution was observed.

EXAMPLES 5.2 TO 5.7

Using the solutions as identified in Table 8 as starting materials, microballoons were prepared in the same manner as in Example 5.1 except that the temperature of the tubular furnace was as shown in Table 8. The evaluation results are shown in Table 9. In each case, the microballoons were spherical polycrystals, and no alkali elution was observed.

Table 8

Example No.	Solute	Concentration (wt%)	Temp. (°C)
5.2	GeCl ₄	1	1200
5.3	SbCl ₃	1	1200
5.4	SnCl ₄	5	1200
5.5	InCl ₃	5	800
5.6	MgCl ₂	5	1500
5.7	CaCO ₃	1	1700

Table 9

Example No.	Crystal phase	Average size (μm)	Bulk density (g/cm ³)	True specific gravity (g/cm ³)
5.2	Tetragonal GeO ₂	2.9	0.024	0.43
5.3	Monoclinic system Sb ₂ O ₅	2.4	0.022	0.39
5.4	Rutile type SnO ₂	2.6	0.019	0.29
5.5	Tetragonal system In ₂ O ₃	2.4	0.039	0.69
5.6	Sodium Chloride type MgO	2.6	0.019	0.29
5.7	Sodium chloride type CaO	2.3	0.019	0.33

EXAMPLE 6.1

A dispersion having 0.63 wt% of aluminum oxide particles having a particle size of 6 nm and 0.37 wt% of silicon dioxide particles having a particle size of 3 nm dispersed in water, was prepared. This dispersion was atomized at a flow rate of 1 ml/min to form droplets having an average particle size of 30 μm, which were introduced into a tubular furnace wherein atmospheric air was maintained at 1400°C. The obtained microballoons were spherical polycrystals of mullite (3Al₂O₃•2SiO₂) having an average particle size of 2.5 μm, a bulk density of 0.023 g/cm³, a true specific gravity of 0.44 g/cm³ and a pressure collapse strength of 2060 kg/cm². No alkali elution was observed.

EXAMPLE 6.2 TO 6.9

Using the dispersions as identified in Table 10 as starting materials, microballoons were prepared in the same manner as in Example 6.1 except that the temperature of the tubular furnace was as shown in Table 10. The evaluation results are shown in Table 11. In each case, the microballoons were spherical polycrystals, and no alkali elution was observed.

Table 10

Example No.	Dispersed particles	Particle size (nm)	Concentration (wt%)	Temp. (°C)
6.2	Iron oxide	6	6.60	1400
	Zinc oxide	3	3.40	
6.3	Aluminium oxide	6	5.61	1700
	Titanium oxide	3	4.39	
6.4	Titanium oxide	3	1.72	1400
	Barium oxide	3	3.28	
6.5	Magnesium oxide	3	1.38	1400
	Aluminium oxide	6	3.49	
	Silicon dioxide	3	5.13	
6.6	Aluminium oxide	6	7.18	1800
	Magnesium oxide	3	2.82	
6.7	Silicon dioxide	6	7.18	1500
	Magnesium oxide	3	2.82	
6.8	Zirconium oxide	4	6.65	1400
	Silicon dioxide	3	3.35	
6.9	Silicon dioxide	6	4.27	1700
	Magnesium oxide	3	5.73	

Table 11

Example No.	Crystal phase	Average size (μm)	Bulk density (g/cm ³)	True specific gravity (g/cm ³)	Pressure collapse strength (kg/cm ²)
6.2	ZnFe ₂ O ₄	1.8	0.022	0.46	1700
6.3	Monoclinic system AlO ₃ • TiO ₂	2.9	0.023	0.41	1360
6.4	Tetragonal system BaTiO ₃	1.6	0.033	0.51	1430
6.5	Cordierite	1.5	0.013	0.23	1560
6.6	Spinel	2.5	0.022	0.38	1760
6.7	Steatite	2.4	0.018	0.38	1560
6.8	Zircon	1.5	0.023	0.49	1760
6.9	Forsterite	2.4	0.022	0.42	1560

EXAMPLE 7.1

Metal iron and barium oxide each having an average particle size of 10 μm, were dissolved in 50 wt% nitric acid to obtain an aqueous nitric acid solution having an iron concentration of 8.3 wt% and a barium oxide concentration of 1.9 wt%. This solution was atomized at a flow rate of 0.8 ml/min to form droplets having an average particle size of 30 μm, which were introduced into a tubular furnace wherein atmospheric air was maintained at 1450 °C.

EXAMPLE 7.2

Zirconium oxychloride octahydrate and calcium carbonate were diluted with water to obtain a solution containing 4.85 wt% of ZrOCl₂ • 8H₂O and 0.25 wt% of CaCO₃. This solution was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of 30 μm, which were introduced into a tubular furnace wherein atmospheric air was maintained at 1600 °C.

EXAMPLE 7.3

Lead chloride and titanium tetrachloride were dissolved in pure water to obtain an aqueous solution having a lead concentration of 4.06 wt% and a titanium concentration of 0.94 wt% as calculated as the respective metals. This solution was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of 30 μm , which were introduced into a tubular furnace wherein atmospheric air was maintained at 1200 °C.

EXAMPLE 7.4

A dispersion having 4.04 wt% of aluminum oxide having a particle size of 6 nm and 4.77 wt% of silicon dioxide particles having a particle size of 3 nm dispersed in an aqueous solution containing 0.952 wt% of lithium hydroxide, was prepared. This dispersion was atomized at a flow rate of 1 ml/min to form droplets having an average particle size of 30 μm , which were introduced into a tubular furnace wherein atmospheric air was maintained at 1200 °C.

EXAMPLE 7.5

A dispersion having 2.74 wt% of aluminum oxide particles having a particle size of 6 nm and 6.46 wt% of silicon dioxide particles having a particle size of 3 nm dispersed in an aqueous solution containing 0.64 wt% of lithium hydroxide, was prepared. This dispersion was atomized at a flow rate of 1 ml/min to form droplets having an average particle size of 30 μm , which were introduced into a tubular furnace wherein atmospheric air was maintained at 1200 °C.

In each of Examples 7.1 to 7.5, the obtained microballoons were spherical polycrystals, and no alkali elution was observed. The evaluation results thereof are shown in Table 12.

Table 12

Example No.	Crystal phase	Average size (μm)	Bulk density (g/cm^3)	True specific gravity (g/cm^3)	Pressure collapse strength (kg/cm^2)
7.1	$\text{BaO} \cdot 6\text{Fe}_2\text{O}_3$	4.3	0.024	0.34	1710
7.2	CaO stabilized ZrO_2	2.6	0.019	0.29	2660
7.3	Tetragonal system PbTiO_3	5.9	0.024	0.40	630
7.4	Eucryptite	1.9	0.013	0.33	1560
7.5	β -spodumen	1.9	0.018	0.39	1560

EXAMPLE 8.1

Metal aluminum powder having an average particle size of 10 μm was dissolved in 10 wt% nitric acid to obtain a 2 wt% aluminum nitrate aqueous solution. This solution was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of 160 μm , which were introduced into a tubular furnace wherein a mixed atmosphere comprising 98% of nitrogen and 2% of hydrogen, was maintained at 1200 °C. The evaluation results of the obtained microballoons are shown in Table 13.

EXAMPLE 8.2

Microballoons were prepared in the same manner as in Example 8.1 except that the average particle size of the droplets was 90 μm . The evaluation results are shown in Table 13.

EXAMPLE 8.3

A dispersion having 0.12 wt% of aluminum oxide particles having a particle size of 3 nm and 0.36 wt% of carbon powder dispersed in water, was prepared. This dispersion was atomized at a flow rate of 1 ml/min to form droplets having an average particle size of 30 μm , which were introduced at a flow rate of 1 ml/min into a two step tubular furnace comprising a tubular furnace wherein a nitrogen atmosphere was maintained at 2200 °C and a tubular furnace wherein a nitrogen atmosphere was maintained at 1200 °C, connected in

series.

EXAMPLE 8.4

5 A dispersion having 1 wt% of silicon having an average particle size of 2 nm dispersed, was prepared. This dispersion was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of 30 μ m, which were introduced into a two step tubular furnace comprising a tubular furnace wherein a nitrogen atmosphere was maintained at 1500 °C and a tubular furnace wherein a nitrogen atmosphere was maintained at 1400 °C, connected in series.

EXAMPLE 8.5

10 An aqueous solution having 2 wt% of tantalum chloride dissolved therein, was prepared. This solution was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of 30 μ m, which were introduced into a two step tubular furnace comprising a tubular furnace wherein an ammonia atmosphere was maintained at 2000 °C and a tubular furnace wherein ammonia an atmosphere was maintained at 1300 °C.

EXAMPLE 8.6

20 A dispersion having 2.8 wt% of metal silicon having a particle size of 2 nm and 3.6 wt% of carbon powder dissolved in water, was prepared. This dispersion was atomized at a flow rate of 1 ml/min to form droplets having an average particle size of 30 μ m, which were introduced into a two step tubular furnace comprising a tubular furnace wherein a carbon-installed vacuum atmosphere was maintained at 2200 °C and a tubular furnace wherein a carbon-installed vacuum atmosphere was maintained at 1800 °C, connected in series.

EXAMPLE 8.7

30 A solution containing 1 wt% of zirconium nitrate was prepared. This solution was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of 30 μ m, which were introduced into a two step tubular furnace comprising a tubular furnace wherein a carbon-installed vacuum atmosphere was maintained at 2200 °C and a tubular furnace wherein a carbon-installed vacuum atmosphere was maintained at 1800 °C, connected in series.

EXAMPLE 8.8

35 A dispersion having 18.3 wt% of metal tungsten having a particle size of 4 nm and 1.2 wt% of carbon powder dispersed in water, was prepared. This dispersion was atomized at a flow rate of 1 ml/min to form droplets having an average particle size of 30 μ m, which were introduced into a two step tubular furnace comprising a tubular furnace wherein a carbon-installed vacuum atmosphere was maintained at 2200 °C and a tubular furnace wherein a carbon-installed vacuum atmosphere was maintained at 1800 °C, connected in series.

EXAMPLE 8.9

45 A dispersion having 1 wt% of silver powder having a particle size of 3 nm dispersed in water, was prepared. This dispersion was atomized at a flow rate of 1 ml/min to form droplets having an average particle size of 30 μ m, which were introduced into a two step tubular furnace comprising a tubular furnace wherein an argon atmosphere was maintained at 950 °C and a tubular furnace wherein an argon atmosphere was maintained at 700 °C, connected in series.

EXAMPLE 8.10

55 A dispersion having copper sulfide having an average particle size of 0.2 μ m and copper oxide having an average particle size of 0.1 μ m dispersed in pure water in a molar ratio of $\text{Cu}_2\text{S}/\text{Cu}_2\text{O} = 0.5$ and at a total concentration of 1 wt%, was prepared. This dispersion was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of 30 μ m, which were introduced into a tubular furnace

wherein an argon atmosphere was maintained at 1100 °C.

EXAMPLE 8.11

- 6 Lead hydrogencarbonate having a particle size of 3 μm was dissolved in a 5 wt% acetic acid solution to obtain a solution containing 1 wt% of $\text{Pb}_3(\text{CO}_3)_2(\text{OH})_2$. This solution was atomized at a flow rate of 0.6 ml/min to form droplets having an average particle size of 30 μm , which were introduced into a two step tubular furnace comprising a tubular furnace wherein a carbon monoxide atmosphere was maintained at 800 °C and a tubular furnace wherein a carbon monoxide atmosphere was maintained at 300 °C, connected
- 10 in series.

In each of Examples 8.1 to 8.11, the microballoons were spherical polycrystals, and no alkali elution was observed. The evaluation results thereof are shown in Table 13.

Table 13

Example No.	Crystal phase	Average size (μm)	Bulk density (g/cm^3)	True specific gravity (g/cm^3)	Pressure collapse strength (kg/cm^2)
8.1	AlN	6.6	0.030	0.41	2200
8.2	AlN	9.3	0.034	0.51	2600
8.3	AlN	1.6	0.044	0.53	2090
8.4	Si_3N_4	1.2	0.034	0.45	2160
8.5	TaN	2.8	0.071	0.99	-*
8.6	SiC	1.2	0.034	0.43	2690
8.7	ZrC	1.9	0.064	0.81	2230
8.8	WC	2.2	0.084	1.03	2090
8.9	Metal silver	1.3	0.069	0.83	690
8.10	Metal copper	3.2	0.074	0.95	560
8.11	Metal lead	2.8	0.081	0.99	-*

* not examined

Claims

1. A process for producing crystalline microballoons, which comprises atomizing a solution or dispersion having an inorganic material or its precursor dissolved or dispersed in a liquid medium, to form droplets of the solution or dispersion, supplying the droplets into a high temperature atmosphere adjusted so that the liquid medium would rapidly evaporate and the inorganic material or its precursor would sinter or fuse to form crystalline microballoons of the inorganic material, and recovering the microballoons thereby formed.
2. The process for producing crystalline microballoons according to Claim 1, wherein the concentration of the inorganic material or its precursor in the solution or dispersion is from 0.1 to 80 wt%.
3. The process for producing crystalline microballoons according to Claim 1 or 2, wherein the droplets have an average particle size of from 0.1 to 1000 μm .
4. The process for producing crystalline microballoons according to Claim 1, 2 or 3, wherein the microballoons have a bulk density of from 0.01 to 2.0 g/cm^3 and an average particle size of from 0.1 to 300 μm .
5. The process for producing crystalline microballoons according to Claim 1, 2, 3 or 4, wherein the liquid medium is water, and the high temperature atmosphere has a temperature of from 300 to 2200 °C.
6. The process for producing crystalline microballoons according to Claim 1, 2, 3, 4 or 5, wherein the inorganic material is a metal oxide, a metal nitride, a metal carbide, or a metal.

7. The process for producing crystalline microballoons according to Claim 1, 2, 3, 4, 5 or 6, wherein the high temperature atmosphere has a temperature not higher than the melting point of the inorganic material.
8. Microballoons made of a sintered polycrystalline material and having spherical hollow structures, which have an average particle size of from 0.1 to 300 μm and a bulk density of from 0.01 to 2.0 g/cm^3 .
9. Microballoons made of a polycrystalline material and having spherical hollow structures having an alkali elution degree of at most 0.01 meq/g, which have an average particle size of from 0.1 to 300 μm and a bulk density of from 0.01 to 2.0 g/cm^3 .
10. The microballoons according to Claim 8 or 9, wherein the polycrystalline material is a metal oxide, a metal nitride, a metal carbide, or a metal.

FIGURE 1

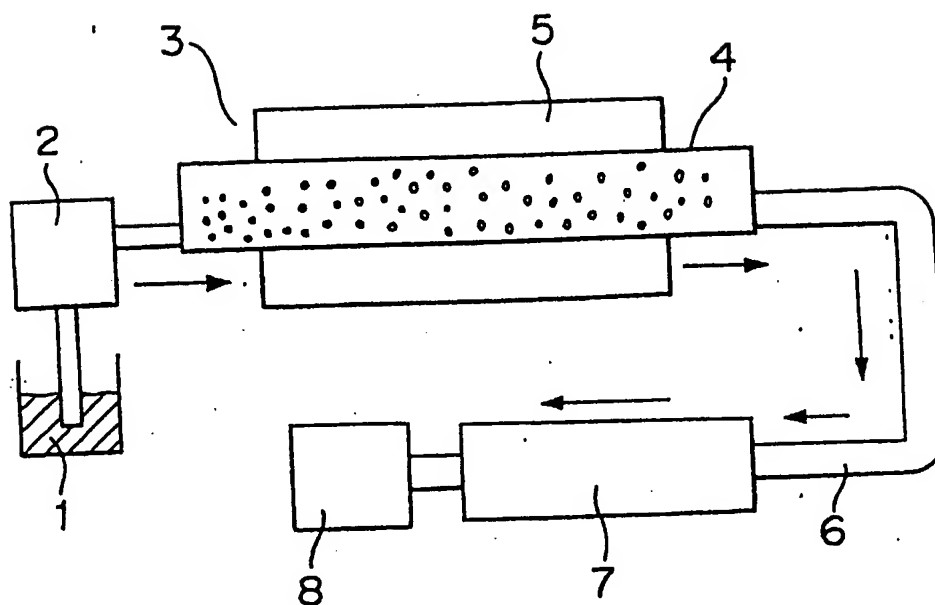
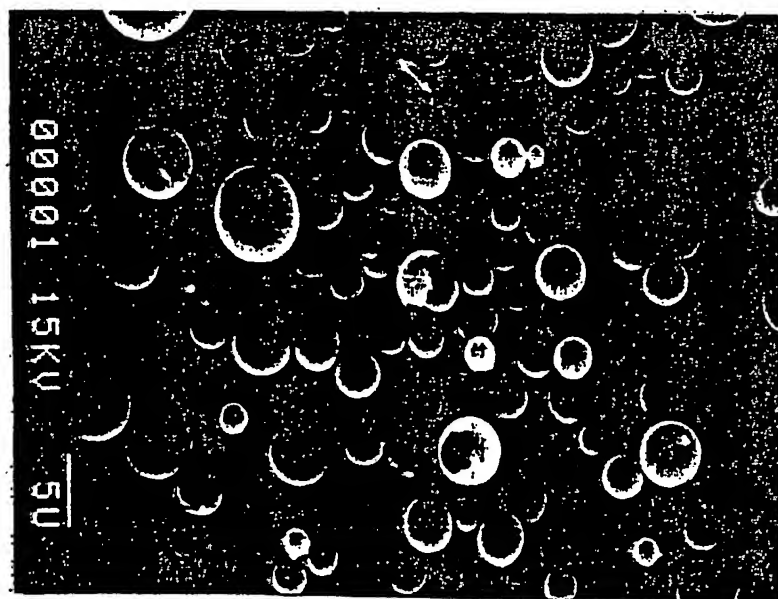


FIGURE 2





European Patent
Office

EUROPEAN SEARCH REPORT

Application Number
EP 93 11 9965

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cls)
X	PATENT ABSTRACTS OF JAPAN vol. 14, no. 242 (C-0721)23 May 1990 & JP-A-02 064 016 (SAKAMOTO MITSUHISA) 5 March 1990 * abstract *	1,3,5-7	B01J13/04
X	PATENT ABSTRACTS OF JAPAN vol. 12, no. 40 (C-474)5 February 1988 & JP-A-62 191 426 (ETSURO KATO) 21 August 1987 * abstract *	1,3,5-7	
X	PATENT ABSTRACTS OF JAPAN vol. 8, no. 71 (C-217)3 April 1984 & JP-A-58 223 606 (NIPPON SODA) 26 December 1983 * abstract *	1,3,5-7	
X	WO-A-92 09543 (MANVILLE CORPORATION) * claims; figures *	1,3,5-7	
A	US-A-4 349 456 (SOWMAN)		TECHNICAL FIELDS SEARCHED (Int.Cls)
A,D	US-A-3 365 315 (BECK ET AL)		B01J
A,D	US-A-5 064 784 (SAITO ET AL.)		C01B
			B22F
			C04B
The present search report has been drawn up for all claims			
Place of search THE HAGUE		Date of completion of the search 9 March 1994	Examiner Meertens, J
CATEGORY OF CITED DOCUMENTS		T : theory or principle underlying the invention E : earlier patent document, but published on, or after the filing date D : document cited in the application L : document cited for other reasons & : member of the same patent family, corresponding document	
X : particularly relevant if taken alone Y : particularly relevant if combined with another document of the same category A : technological background O : non-written disclosure P : intermediate document			

EPF FORM 1500 (04/93) (P.O. 1/93)